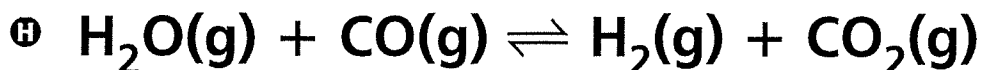
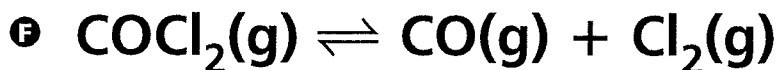
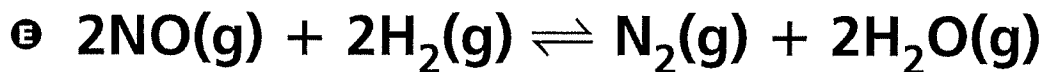
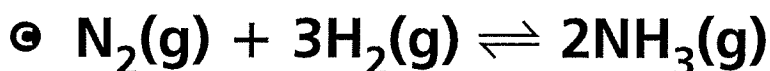
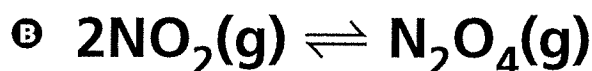
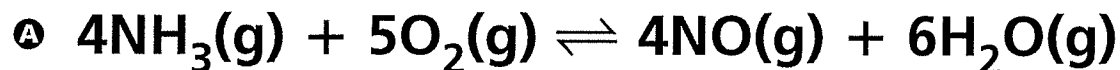


Chemical Equilibrium

Use with Chapter 17,
Section 17.1

TEACHING TRANSPARENCY WORKSHEET



Chemical Equilibrium

Use with Chapter 17,
Section 17.1

1. Write the equilibrium constant expressions for the reactions on the transparency.

A _____ E _____

B _____ F _____

C _____ G _____

D _____ H _____

2. What is heterogeneous equilibrium?

3. Which equilibrium reactions are homogeneous equilibria?

4. Which equilibrium reaction is a heterogeneous equilibrium?

5. When writing equilibrium constant expressions, pure solids and liquids are not included. Why? Why do you include all reactants of equation D in its equilibrium constant expression?

6. If the K_{eq} for one of the reactions was 35.6, what would you know about the equilibrium?

