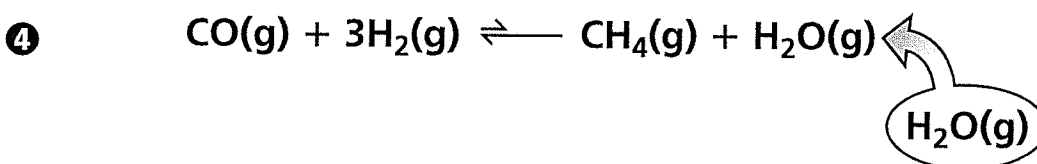
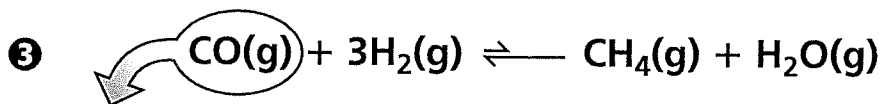
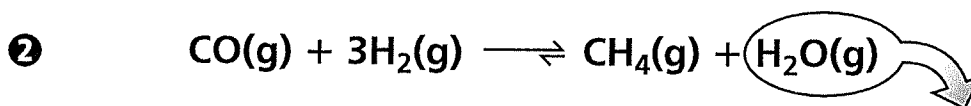
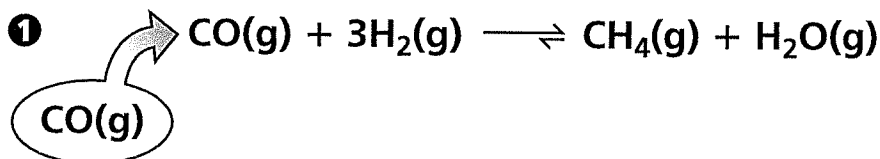
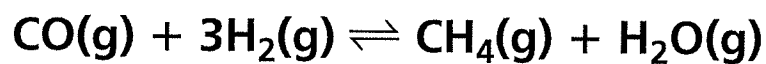


## TEACHING TRANSPARENCY MASTER

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# How Changing Concentration Affects Equilibrium

Use with Chapter 17,  
Section 17.2

**TEACHING TRANSPARENCY WORKSHEET**

# How Changing Concentration Affects Equilibrium

Use with Chapter 17,  
Section 17.2

1. Why do the changes shown in equations 1 and 2 cause the equilibrium to move to the right? What other changes in concentration would also cause a shift to the right?

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2. Why do the changes shown in equations 3 and 4 cause the equilibrium to move to the left? What other changes in concentration would also cause a shift to the left?

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3. What effect would decreasing the volume of the reaction container have on the equilibrium? Why?

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4. When does changing the volume of the reaction container not affect the equilibrium?

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5. The production of methane and water vapor from carbon monoxide and hydrogen gas is an exothermic reaction. What does this tell you about how an increase in temperature would affect the equilibrium of this reaction? How would it affect the equilibrium constant?

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6. When does changing the temperature not affect a reaction at equilibrium?

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