



## Le Chatelier's Principle: Webquest

---

Go to the sites shown here.

<https://ed.ted.com/on/moKExh1B>

<https://courses.lumenlearning.com/boundless-chemistry/chapter/factors-that-affect-chemical-equilibrium/>

See additional resources on

<http://vcichemistry.weebly.com/simulations.html>

Answer the questions based on what you've read.

1. What does LeChatelier's principle mean?
2. What is a stress?
3. If you ADD something (a reactant, product or heat/energy), will the equilibrium shift toward the side of the reaction to make even more of it, or will the equilibrium shift in the direction to use it up? Explain.
4. If you REMOVE a chemical or heat from a system, will the system shift toward the side that replaces what you took out or try to use even more of it?
5. Explain what happens when you increase pressure on a system that was at equilibrium?
6. View this Ted Ed video on an important application of Le Chat's Principle (<https://ed.ted.com/on/nLHJykrk>). Write down what Haber did to optimize the production of ammonia and how it relates to the Le Chat's principle.

Additional resources:

<http://www.chem1.com/acad/webtext/chemeq/Eq-02.html>

<http://www.chem.ox.ac.uk/vrchemistry/ChemicalEquilibrium/HTML/page29.htm>

<http://www.youtube.com/watch?v=4-fEvVNTIE>